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sentence or statement is true or
false. T F. 1. Nuclear changes
generally absorb more energy
than chemical changes. T F. 2. In

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4.18 J (1.00 g) gC qmcT=Δ = .°

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(10.5 °C) $q = 43.9 \text{ J}$

Statement: The amount of energy absorbed by 1.00 g of water, $\text{H}_2\text{O}(\text{l})$, when its temperature is raised 10.5 °C is 43.9 J. 36. Given: $V \text{ HCl}(\text{aq}) = 100.0 \text{ mL}$; $V \text{ NaOH}(\text{aq}) = 100.0 \text{ mL}$; $T_{\text{initial}} = 23.5 \text{ °C}$; T

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Chapter 5 Review, pages 338-339 - Pre University Courses

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Chapter 5 Quiz. True/False.
Indicate whether the sentence or
statement is true or false. 1.

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Nuclear changes generally absorb more energy than chemical changes. 2. In exothermic reactions, the reactants have more kinetic energy than the products. 3.

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5 Review Answers

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Chapter 1, Organic Compounds,

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Solutions introduces the historical concept of organic compounds as those compounds produced by living things. Students then learn that the current concept of organic chemistry is the study of carbon compounds. Each section of the chapter addresses a different

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group of organic compounds,
related by structure.

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Section 6.5: Rate Law. Tutorial 1
Practice, page 380 1. (a) The order
of reaction with respect to A is 2.

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When the initial concentration of A is doubled, the rate increases by a factor of 4, or 2², so, the exponent of the concentration of reactant A is 2. The order of reaction with respect to B is 1. When the initial concentration of B is doubled, the rate is doubled,

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or multiplied by 21, so the exponent of the concentration of reactant B is 1.

Section 6.5: Rate Law - Pre University Courses

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false. T F. 1. Nuclear changes generally absorb more energy than chemical changes. T F. 2. In exothermic reactions, the reactants have more kinetic energy than the products.
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20.5 °C to 26.0 °C is an exothermic reaction. (d) The dissolution of a salt, which changes the temperature of the solution from 21 °C to 19 °C is an endothermic reaction. Nelson Chemistry 12 Chapter 5 Solutions 30.

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